LAMPIRAN

Lampiran 1 Perhitungan Pembuatan Amilum 1%

Amilum 1% dalam 100 ml

$$\% \text{ Amilum} = \frac{\text{Massa amilum}}{\text{Volume larutan}} \text{ x } 100\%$$

$$1 \text{ \% Amilum} = \frac{\text{Massa amilum}}{100 \text{ ml}} \text{ x } 100\%$$

$$\text{Massa amilum} = \frac{1\% \text{ x } 100 \text{ ml}}{100\%}$$

Massa amilum = 1 gram

Lampiran 2 Perhitungan Pembuatan Kalium Iodida 10%

Kalium iodida 10% dalam 100 ml

$$\% \text{ Kalium iodida} = \frac{\text{Massa Kalium iodida}}{\text{Volume larutan}} \text{ x } 100\%$$

$$10 \text{ \% Kalium iodida} = \frac{\text{Massa Kalium iodida}}{100 \text{ ml}} \text{ x } 100\%$$

$$\text{Massa Kalium iodida} = \frac{10\% \text{ x } 100 \text{ ml}}{100\%}$$

Massa Kalium iodida = 10 gram

Lampiran 3 Perhitungan Pembuatan Perak Nitrat 5%

Larutan perak nitrat 5% dalam 100 ml

$$\% \text{ Perak nitrat} = \frac{\text{Massa Perak nitrat}}{\text{Volume larutan}} \text{ x } 100\%$$

$$5 \text{ \% Perak nitrat} = \frac{\text{Massa Perak nitrat}}{100 \text{ ml}} \text{ x } 100\%$$

$$\text{Massa Perak nitrat} = \frac{5\% \text{ x } 100 \text{ ml}}{100\%}$$

Massa Perak nitrat = 5 gram

Lampiran 4 Perhitungan Pembuatan Asam Nitrat 3N

Asam nitrat 3N dalam 100 ml

BM Asam nitrat = 1 + 14 + 16 x 3

= 15 + 48

= 63 g/mol

BE Asam nitrat = BM x Valensi

= 63 g/mol x 1

$$= 63 \frac{\text{g}}{\text{mol.ek}}$$

$$\text{Normalitas} = \frac{1000 \text{ x BJ x C}}{\text{BE x } 100}$$

$$\text{Normalitas} = \frac{1000 \text{ x } 1{,}42\frac{\text{g}}{\text{ml}} \text{x } 69{,}5\%}{63\frac{\text{g}}{\text{mol.ek}} \text{ x } 100}$$

$$\text{Normalitas} = \frac{98{,}690}{6{,}300}$$

$$\text{Normalitas} = 15{,}6$$

Sehingga :

$$\text{N1 x V1} = \text{N2 x V2}$$

$$15{,}6 \text{ N x V1} = 3\text{N x } 100 \text{ ml}$$

$$\text{V1} = \frac{300\frac{\text{ml}}{\text{N}}}{15{,}6 \text{ N}}$$

$$\text{V1} = 19{,}2 \text{ ml dalam } 100 \text{ ml}$$

Lampiran 5 Dokumentasi Sampel Beras



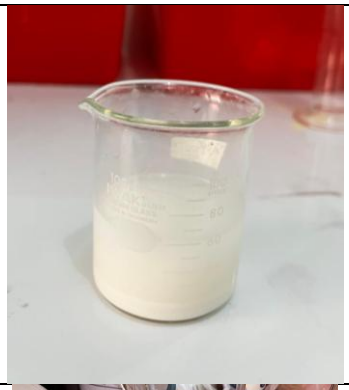
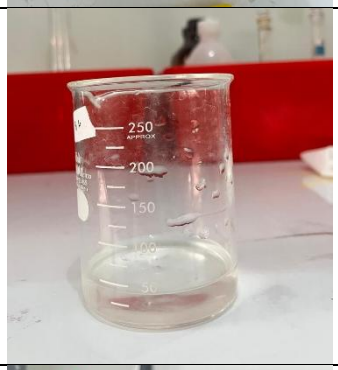
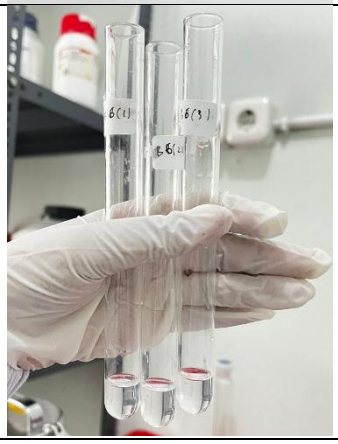
Lampiran 6 Dokumentasi Pemutih

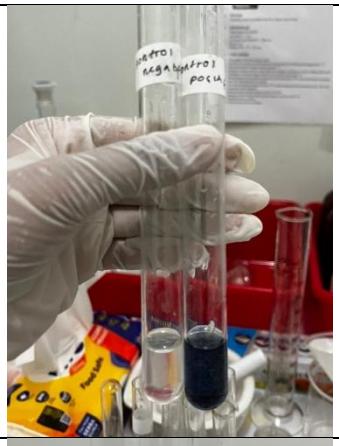
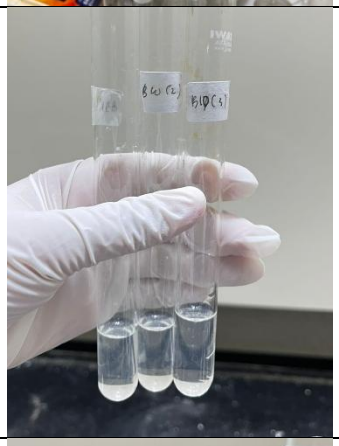


Lampiran 7 Dokumentasi Penelitian

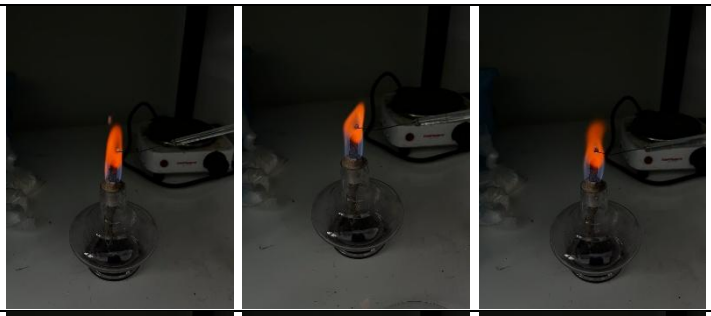
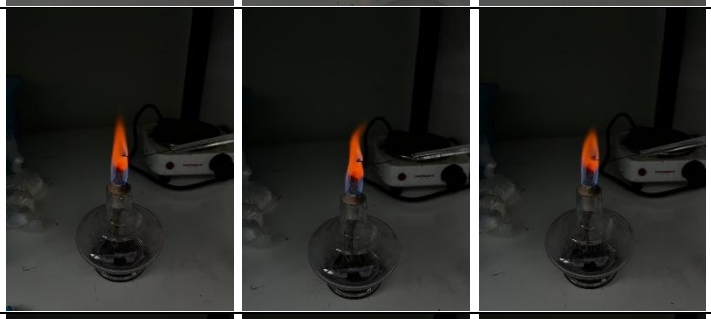
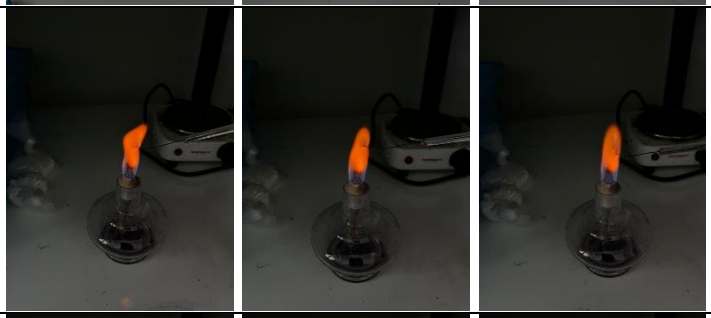
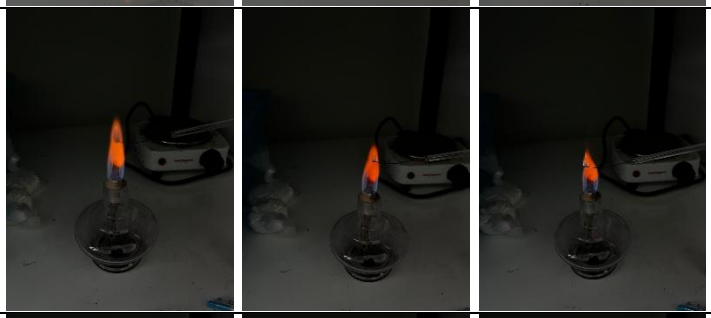
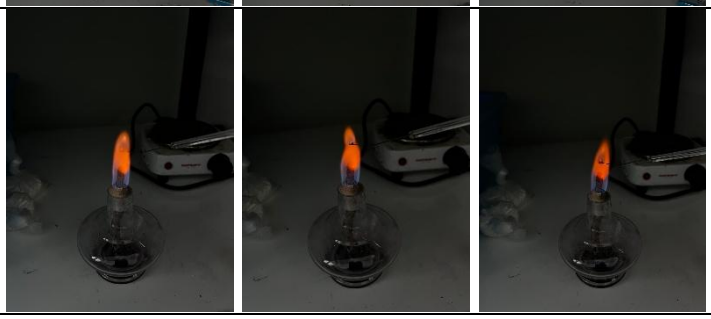
No	Gambar	Keterangan
1.		Penimbangan untuk pembuatan larutan amilum 1%
2.		Penimbangan untuk pembuatan larutan kalium iodide 10%
3.		Penimbangan untuk pembuatan larutan perak nitrat 5%
4.		Membuat larutan asam nitrat 3N

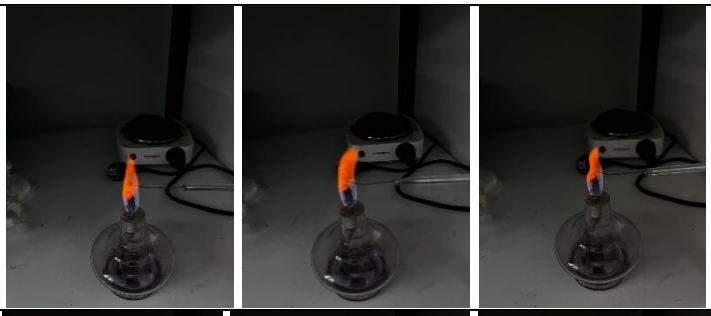
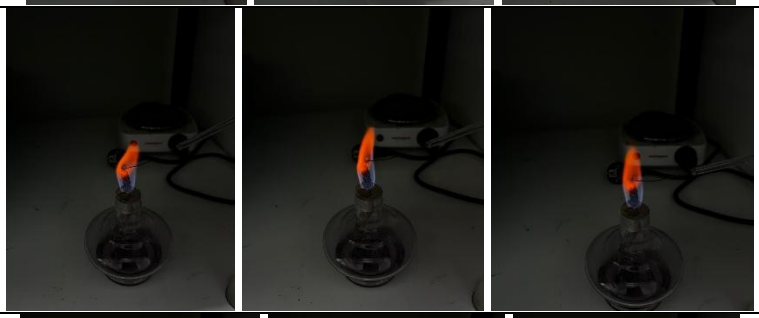
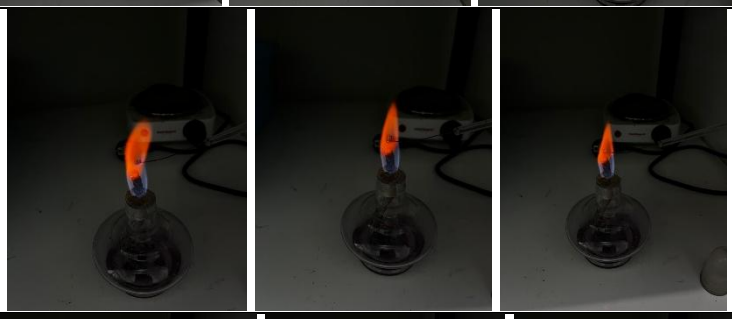
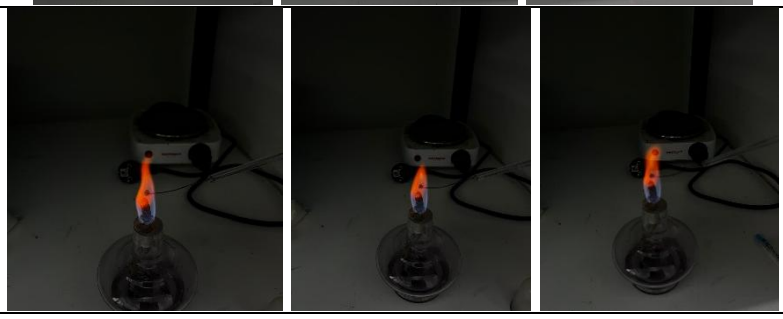
5.		Menghaluskan sampel beras putih
6.		Penimbangan sampel serbuk beras putih
7.		Melakukan pengujian uji nyala api
8.		Penambahan akuadest dalam sampel serbuk beras putih

9.		Diaduk
10.		Penyaringan sampel
11.		Diperoleh hasil filtrat
12.		Uji reaksi warna, Filtrat 2 ml + 5 tetes Kalium Iodida

13.		Uji reaksi warna, Setelah penambahan amilum
14.		Uji pengendapan, filtrat 3 mL + 1 mL asam nitrat
15.		Pemanasan setelah penambahan asam nitrat
16.		Uji pengendapan, setelah penambahan perak nitrat

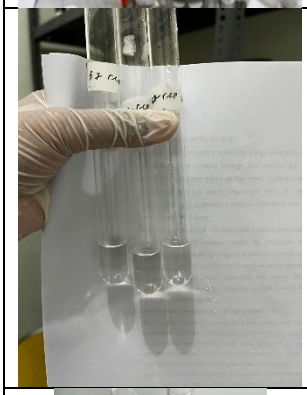
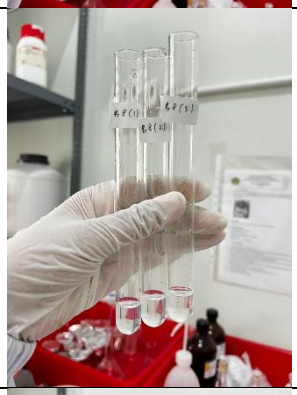
Lampiran 8 Dokumentasi Hasil Uji Uji Nyala

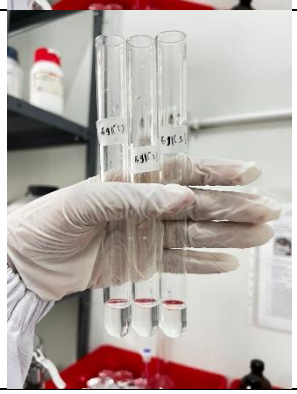
Gambar	Keterangan
	Sampel B1
	Sampel B2
	Sampel B3
	Sampel B4
	Sampel B5

	Sampel B6
	Sampel B7
	Sampel B8
	Sampel B9
	Sampel B10

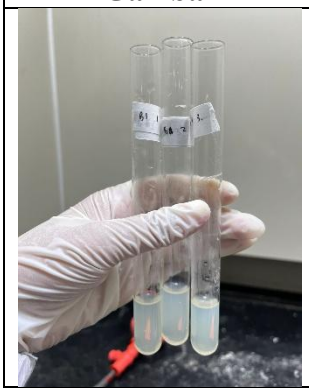
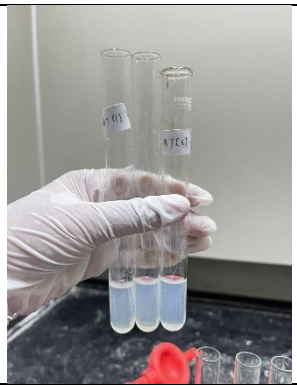
			Sampel B11

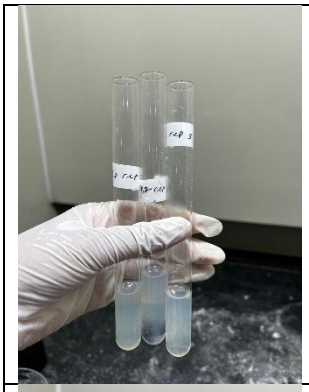
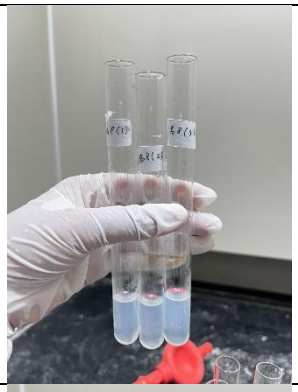
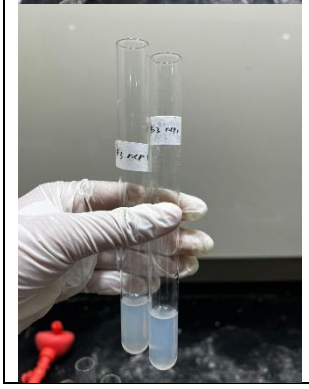
Lampiran 9 Dokumentasi Hasil Uji Reaksi warna

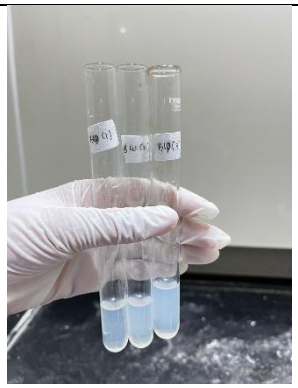
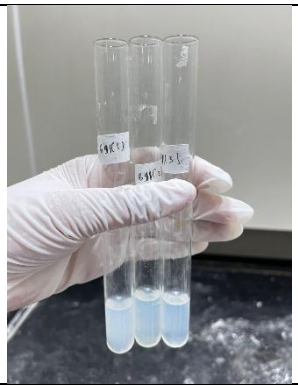
Gambar	Keterangan	Gambar	Keterangan
	Sampel B1		Sampel B7
	Sampel B2		Sampel B8
	Sampel B3		Sampel B9

	Sampel B4		Sampel B10
	Sampel B5		Sampel B11
	Sampel B6		

Lampiran 10 Dokumentasi Hasil Uji Pengendapan

Gambar	Keterangan	Gambar	Keterangan
	Sampel B1		Sampel B7

	Sampel B2		Sampel B8
 	Sampel B3		Sampel B9

 	Sampel B4		Sampel B10
	Sampel B5		Sampel B11
	Sampel B6		